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AP Chemistry 2009 Q&A - College Board

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$\text{CH}_4(\text{g}) + 2 \text{Cl}_2(\text{g}) \rightarrow \text{CH}_2\text{Cl}_2(\text{g}) + 2 \text{HCl}(\text{g})$  3. Methane gas reacts with chlorine gas to form dichloromethane and hydrogen chloride, as represented by the equation above. (a)

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A 25.0 g sample of methane gas is placed in a reaction vessel containing 2.58 mol of  $\text{Cl}_2(\text{g})$ .

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If the temperature is increased for an endothermic reaction ( $\Delta H_{298D} = +131 \text{ kJ mol}^{-1}$ ), then by Le Chatelier's principle the reaction will shift toward products, thereby absorbing energy. With greater concentrations of products at equilibrium, the value of  $K_p$  will increase.

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## AP Chemistry Free Response Question 2009 #5

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